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THIS IS UNEVALUATED INFORMATION FOR THE RESEARCH USE OF TRAINED INTELLIGENCE ANALYSTS

SOURCE Decumentary as indicated. (Information specifically requested,)

RECENTIA PUBLISHED RESEARCH OF THE LENINGRAD STATE PEDIATRIC MEDICINE INSTITUE, USSR

"Interaction of Organomercury Compounds with Halogon Derivatives of Hydrocarbons," M. H. Koton, T. M. Zorina and E. G. Osberg, Leningrad State Ped Med Inst

"Zhurnal Obahchey Khimii" Vol 17, 1947, pp 59-62

Phylig does not react at 130° with either CHCl<sub>3</sub> or CHBr<sub>3</sub> but does react with CHl<sub>2</sub>: a mixture of Ph<sub>2</sub>HG with CHl<sub>3</sub> heated in a scaled tube at 130° 3 hours gave PhHgCl<sub>3</sub> and PhHgI; heating of the same stare at the same temperature five hours gave C<sub>2</sub>I, axi PhHgI. The reaction evidently proceeds in two stages: Ph<sub>2</sub>Hg+CHl<sub>3</sub> → C<sub>6</sub>H<sub>6</sub>+FhHgCl<sub>3</sub>, and 2PhH<sub>2</sub>Cl<sub>3</sub> → 2PhHgI+C<sub>2</sub>I<sub>4</sub>. Hg(OAc)<sub>2</sub> reacts readily with HsI and EtI. Reactions of Hg(OAc)<sub>2</sub> with PhCH<sub>2</sub>Cl, PhBr, CHCl<sub>3</sub>, CHBr<sub>3</sub>, CHI<sub>3</sub>, S<sub>2</sub>H<sub>2</sub>Br<sub>2</sub>, C<sub>2</sub>H<sub>2</sub>Cl<sub>2</sub>, C<sub>2</sub>H<sub>2</sub>Cl<sub>4</sub>, and CCl<sub>4</sub> and their products are described.

"Reaction of Organomercury Compounds of the General Formula RHgX with Phenols," E. M. Koton and T. M. Zorina, Lemingrad State Ped Med Inst

"Zhurnal Obshchey Khimii" Vol 17, 1947, pp 1303-6

Generally, compounds of the type RigX on reacting with pyrogallol break down into RH, Rg212, and oxidation products of pyrogallol. It is possible that the reaction proceeds through the initial formation of free radicals R, which undergo oxidation—

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with EtOH as solvent, are described.

reduction reactions with pyrogallol to give RH and oxidation products of pyrogallol. PhHgI was exceptional in that it gave metallic Hg instead of HgI. The reactions which were run in seeled tubes at 130°

"Reaction of Aromatic Compounds of Tin, Lead, and Biamuth with Phenols," N. W. Roton, Leningrad State Ped Med Inst

"Zhurnal Obshchey Khimii" Vol 17, 1947, pp 1307-8

Metallo-organic compounds, phenols, and alcohol were allowed to react at 130°. Bi derivatives were most reactive, followed by Hg, Pb, and Sn. Ph.Bi and 1-C10H70H, 3 hours, no solvent, gave 98.1% Bi; (C10H7)3BI gave 89.11% Bi; Ph.Pb gave 28.71% Pb; Ph.Sn gave 0% Sn. Lith pyrogallol as the phenol and 200H as solvent, Ph.Bi gave 92.42% Bi in 0.5 hour, 98.03% Bi in 1 hour; (C10H7)3BI gave 17.51% Bi in 1 hour, 49.9% in 2 hours, 64.97% in 3 hours, and 70.5% in 4 hours. Ph.Pb gave 0% Pb in 3 hours and Ph.Sn 0% Sn in 5 hours. The organic reaction products were C6H6 or C10H6, respectively.

"Influence of Sulfonamides on the Protozoan Fauna of Frog Intestines," K. A. Meshchorskaya-Shtoynberg, Lemingrad State Ped Med Inst

"Farmakol i Toksikol" Vol 9, No 4, 1946, pp 36-9

Sulfadiazine (I), sulfathiazolc (II), sulfidine (III), marfamil (para-sulfonamidobenzylamine, IV), sulfanil-amide (V), disulfan (VI), and acriquine (VII) were toxic, in descending order as listed, to: Opalina ranarum, Mictotherus, Mematoda apectana aculeata, Gorgodera gorgoderina, Diplodiscus subclavitus, and Polystomum interrebium. Micotinic acid protects the intestinal protozoa. The observed changes in protozoa can be used to advantage in teaching students the effects of sulfa drugs.

"Comparison of the Antimuscarinic Effects Produced by the Optical Isomers of Comphor," A. M. Rusanov, Leningrad State Ped Med Inst

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"Fisiologicheskiy Zharral SSSR" Vol 32, 1946, pp 283-6

The entimuscarinic effects of the isomers of camphor were studied by subcutaneously injecting a free with solutions of arecoline and carbocholine. When the activity of the heart was arrested, an aqueous solution of camphor was injected subcutaneously and the time until resumption of heart activity was measured. Contrary to other reports, this investigation showed a marked difference in the capacity of the 2 isomers to restore the cardiac activity. The disomer was twice as rapid in its action as the 1-form,

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and was five times as actios.

"Catalytic Decomposition of Acetale," N. M. Koton and L. I. Barsukova, Lemingrad State Ped Med Inst

"Zhurnal Coshchey Khimdi" Vol 16, 1946, pp 685-94

The catalytic decomposition of 3 acetals /MeCH(OBu)2, isovaleraldohyde diisosmyl acetal, and NepCHCH-(OCHZHH42)2/was studied over Cu, Cu-2r, Cu-Ur, and the catelysts in the range 200-500°. In contrast with the behavior of alcohols, the acetals decompose in such a manner as to yield an alcohol, CuHzm, and gaseous products. The olefin is formed apparently because of intermediate formation of vinyl ethers.

11 3 acetals are very stable thermally in the absence of the catalysts.

"Catalytic Decomposition of Formaldehyde Dimethyl Acetal and Formaldehyde Discamyl Acetal," L. M. Koton and I. A. Chernov, Leningrad State Ped Red Inst

"Zhurnel Coshchey Khimii" Vol 16, 1946, pp 695-700

Mathylal is readily decomposed over metallic catalysts with formation of H, CO, CO2, and CH1. Formaldehyde discamyl acetal was more stable and only at 350-5000 (over Cu-Ur or Cu-Zr) gave appreciable amounts of olefins, H, and gaseous decomposition products.

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